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simple tri-atomic molecule, H₂O.

- Each O-H bond is highly polar, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values)
- bond angle of water = 105°.

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Aqueous Systems. surface tension.
surfactant. aqueous solution. solvent.
the inward force or pull that tends to
minimize the surface ar.... any
substance that interferes with
hydrogen bonding between wa.... is
water that contains dissolved

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substances. the dissolving medium in
a solution.

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- Page 501: 17 Answer CH₄ doesn't
dissolve in water because it is a
molecular compound with no net
dipole KCl is soluble because of its
ionic bonds.

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Collect and Organize Figure
P15 shows four lines to describe the
possible dependence of percent
ionization of acetic acid with
concentration. We are to choose the
one that best represents the trend for

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this weak acid. Analyze

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